

Available online at www.sciencedirect.com



Tetrahedron Letters

Tetrahedron Letters 49 (2008) 755–759

## Highly enantioselective Rh-catalyzed hydrogenations with heterocombinations of pentafluorobenzyl- and methoxybenzyl-derived binaphthyl phosphites

Benita Lynikaite<sup>a</sup>, Ján Cvengroš<sup>a</sup>, Umberto Piarulli<sup>b,\*</sup>, Cesare Gennari<sup>a,\*</sup>

<sup>a</sup> Dipartimento di Chimica Organica e Industriale, Centro di Eccellenza C.I.S.I., Università degli Studi di Milano, Via G. Venezian, 21, 20133 Milano, Italy <sup>b</sup> Dipartimento di Scienze Chimiche e Ambientali, Università degli Studi dell'Insubria, Via Valleggio, 11, 22100 Como, Italy

> Received 20 November 2007; accepted 30 November 2007 Available online 4 December 2007

## Abstract

The Rh-catalyzed hydrogenations of dimethyl itaconate and methyl acetamido acrylate using selected heterocombinations of pentafluorobenzyl- and methoxybenzyl-derived binaphthyl phosphites proved to be highly enantioselective (ee 93–99%). In these selected cases the Rh-heterocomplexes, which were formed in a statistical amount (ca. 50% by 31P NMR), turned out to be more active and selective than the two homocomplexes.

 $© 2007 Elsevier Ltd. All rights reserved.$ 

Keywords: Phosphites; Asymmetric hydrogenation; Rhodium; Enantioselective catalysis

In recent years, monodentate phosphorus ligands have held the stage in asymmetric catalysis.<sup>[1](#page-3-0)</sup> In addition to their outstanding activity and selectivity, comparable or even superior to those of bidentate ligands, their convenient, fast and practical synthesis from commercially available mate-rials underlines their potential for industrial applications.<sup>[2](#page-3-0)</sup> Furthermore, the simple preparation of the phosphite, $3$ phosphonite<sup>[4](#page-3-0)</sup> and phosphoramidite<sup>[5](#page-3-0)</sup> functional groups, which are the most frequently represented in monodentate ligands, allows for a facile variation of their design.

Still, the presence of a bidentate ligand is sometimes an inevitable requirement for a successful (stereo)chemical result. Supramolecular bidentate ligands, based on the self-assembly of monodentate ligands possessing complementary functionalities, thus represent a clever solution which combines the important features of bidentate ligands (preorganization and rigidity) with the economy and effi-ciency of monodentate ligands.<sup>[6](#page-3-0)</sup>

The non-covalent assembly of monodentate ligands reported so far in the literature relies mostly on coordinative bonding of nitrogen to zinc<sup>[7](#page-3-0)</sup> or hydrogen bonding.<sup>[8](#page-3-0)</sup> More recently, ionic interactions<sup>[9](#page-3-0)</sup> and the formation of inclusion complexes<sup>[10](#page-3-0)</sup> were exploited to bring monodentate ligands into close proximity. We felt intrigued by these approaches and decided to investigate other non-covalent interactions, that is, the perfluroarene-arene  $\pi-\pi$  interactions.

Perfluroarene-arene  $\pi-\pi$  interactions were observed for the first time in  $1960$ .<sup>[11](#page-3-0)</sup> Their relevance in the solid state has been confirmed several times since then, $12$  and they have found widespread application in supramolecular design and crystal engineering.[13](#page-3-0) The contribution of perfluoroarenearene  $\pi-\pi$  interactions to the efficient photodimerization of olefins in the solid state was also investigated.<sup>[14](#page-3-0)</sup> On the contrary, only a few papers report a quantitative description of the perfluoroarene-arene  $\pi-\pi$  interactions in solution,[15](#page-3-0) and a few applications supposedly exploited them in solution-phase synthetic processes.[16](#page-4-0)

<sup>\*</sup> Corresponding authors. Tel.: +39 025031 4091; fax: +39 025031 4072 (C.G.); tel.: +39 031238 6444; fax: +39 031238 6449 (U.P.).

E-mail addresses: [umberto.piarulli@uninsubria.it](mailto:umberto.piarulli@uninsubria.it) (U. Piarulli), [cesare.](mailto:cesare. gennari@unimi.it) [gennari@unimi.it](mailto:cesare. gennari@unimi.it) (C. Gennari).

<sup>0040-4039/\$ -</sup> see front matter © 2007 Elsevier Ltd. All rights reserved. doi:10.1016/j.tetlet.2007.11.201

In this Letter we report the synthesis of a series of novel perfluroarene- and methoxyarene-derived phosphites, which were used to investigate (by  ${}^{31}P$  NMR) the possible formation of supramolecular bidentate ligands in the Rh-complexes via a non-covalent assembly induced by the perfluroarene–methoxyarene  $\pi-\pi$  interactions. Homo- and hetero-combinations of these ligands were used in enantioselective Rh-catalyzed hydrogenations of dimethyl itaconate and methyl acetamido acrylate.

Two strategies were applied to the synthesis of the designed binaphthyl phosphites (Scheme 1). The commercially available methoxy substituted phenols were refluxed with  $\text{PC1}_3$  to give the corresponding methoxy substitutedphenyl dichlorophosphites (Route A). Subsequent reaction with BINOL in the presence of  $Et_3N$  furnished the electron-rich phosphites  $(R)-1-(R)-3$  (Fig. 1), which were obtained in moderate yields as pure compounds after chromatographic purification.

In the case of methoxy substituted benzylalcohols, an alternative approach (Route B) was followed. BINOL was thus treated with  $\text{PCl}_3$  in the presence of a catalytic amount of  $NMP<sup>17</sup>$  $NMP<sup>17</sup>$  $NMP<sup>17</sup>$  to give the corresponding binaphthyl chlorophosphite, which was then reacted with benzylalcohols in the presence of  $Et_3N$  to afford phosphites  $(R)-4-(R)-8$ (Fig. 1). The same protocol was used also for the synthesis



Scheme 1. Synthesis of benzylalcohol- and phenol-derived binaphthyl phosphites.

of the electron-poor phosphites  $(R)$ -9 and  $(S)$ -10, containing the pentafluorobenzyl moiety.

We thus synthesized a small library of eight electron-rich phosphites (including six BINOL and two octahydro-BINOL derivatives) and two electron-poor pentafluorobenzyl binaphthyl phosphites (Fig. 1).

With the ligand library in our hands, we proceeded with the 31P NMR investigations of perfluoroarene–methoxyarene  $\pi-\pi$  interactions in the rhodium complexes. At first, the  $31P$  NMR spectra of a single ligand in the presence of Rh(acac)(eth)<sub>2</sub> (L:Rh ratio = 2:1) were recorded revealing the fast formation of the corresponding homocomplexes  $L_2Rh(acac)$  ([Fig. 2](#page-2-0)). Then two different ligands (one electron-rich and the other electron-poor) were mixed in the presence of the Rh source (ratio 1:1:1) and the formation of two homocomplexes and one heterocomplex was observed. The NMR spectra were recorded in a broad range of solvents (CDCl<sub>3</sub>, THF- $d_8$ , CD<sub>3</sub>OD, *i*-PrOH/CD<sub>3</sub>OD,  $CD_2Cl_2$ , toluene- $d_8$ , CCl<sub>4</sub>). In all these solvents, the observed ratio was essentially statistical within the experimental error  $(RhL<sub>A</sub>L<sub>A</sub>:RhL<sub>B</sub>L<sub>B</sub>:RhL<sub>A</sub>L<sub>B</sub> = 1:1:2), which means that$ there is no detectable preference for the formation of the heterocomplex due to the methoxyarene–perfluoroarene interactions. In crystals featuring arene–perfluoroarene interactions, the packing energy is dispersion-dominated and coulombic terms are selective rather than quantitatively predominant.<sup>12f</sup> This means that these interactions are likely to be very weak in solution.<sup>15b,c</sup>

Although we did not observe any relevant  $\pi-\pi$ -interaction between electron-rich and electron-poor ligands, we screened the library, using both homo- and heterocombinations of the ligands, in the Rh-catalyzed asymmetric hydrogenation of prochiral alkenes. Typically, the substrate was added to a stirred mixture of  $Rh(cod)BF_4$  (1 mol %) and ligand  $[2 \text{ mol } \%$  (1:1 ratio in the case of heterocombinations)] in dichloromethane and the mixture was hydrogenated for 24 h at room temperature and pressure.

In the Rh-catalyzed asymmetric hydrogenation of dimethyl itaconate, the use of homocombinations of



Fig. 1. Library of phosphite ligands (isolated yields are reported in the brackets).

<span id="page-2-0"></span>

Fig. 2. Studies of perfluoroarene–alkoxyarene  $\pi-\pi$  interactions in the rhodium complexes by <sup>31</sup>P NMR spectroscopy.

monodentate ligands gave the reaction product in high yields (98–100%, Table 1). Methoxyphenol derived ligands  $(R)$ -1,  $(R)$ -2 and  $(R)$ -3 gave the hydrogenated product in high enantiomeric excesses ( $\geq$ 97% ee, entries 1–3). A product of lower enantiomeric purity was obtained with the methoxybenzylalcohol derived ligands  $(R)$ -4– $(R)$ -6 (70– 84% ee, entries 4–6). Only modest ee's were obtained with the octahydro-BINOL derivatives  $(R)$ -7 and  $(R)$ -8 (50–54%) ee, entries 7 and 8). The pentafluorobenzylalcohol derived phosphite  $(R)$ -9 [or  $(S)$ -10] afforded the product in 92% ee (entries 9 and 10).

The ligand heterocombinations were then studied, and also in these cases a complete conversion was observed in most of the reactions. Enhanced activity and selectivity were observed combining ligands  $(R)$ -4 and  $(R)$ -9 [84% ee for  $(R)$ -4, 92% ee for  $(R)$ -9 and 99% ee for the heterocombination of  $(R)$ -4 and  $(R)$ -9, entry 17].

A similar trend was observed in the Rh(I)-catalyzed asymmetric hydrogenation of methyl acetamido acrylate with selected ligands ([Table 2](#page-3-0)). Complete conversions and modest enantiomeric excesses (74–88%) were achieved in all cases. The heterocombination of ligands  $(R)$ -6 and  $(R)$ -9 was the only case where a significantly higher enantiomeric excess (93%, entry 13) was obtained in comparison to the corresponding homocombinations.

In conclusion, we prepared a library of electron-rich phosphites (including BINOL and octahydro-BINOL derivatives) and one electron-poor pentafluorobenzyl binaphthyl phosphite. Although no preferential formation of the Rh-heterocomplexes was observed by  $^{31}P$  NMR

Table 1

Rh(I)-catalyzed asymmetric hydrogenations of dimethyl itaconate

$H_3$ COOC $\bigcap$ COOCH <sub>3</sub> $\frac{Rh(cod)_2BF_4/Ligand(s)}{H_2(1 atm), DCM, rt, 24 h}$ $H_3COOC\bigcap_{*}COOCH_3$	



<span id="page-3-0"></span>Table 2 Rh(I)-catalyzed asymmetric hydrogenation of methyl acetamido acrylate

$Rh(cod)2BF4$ / Ligand(s) H <sub>2</sub> (1 atm), DCM, rt, 24 h										
Homocombinations			Heterocombinations							
Entry	Ligand	Conv. $(\%)$	ee $(\% )$ (abs. conf.)	Entry	Ligands	Conv. $(\%)$	ee $(\% )$ (abs. conf.)			
1	$(R)$ -1	100	86 (S)	8	$(R) - 1/(R) - 9$	100	84 (S)			
2	$(R) - 2$	100	86 (S)	9	$(R) - 2/(R) - 9$	100	86 (S)			
3	$(R) - 3$	100	82(S)	10	$(R) - 3/(R) - 9$	100	86 (S)			
4	$(R) - 4$	100	88 (S)	11	$(R) - 4/(R) - 9$	100	86 (S)			
5	$(R)$ -5	100	84 (S)	12	$(R)$ -5/ $(R)$ -9	100	88 (S)			
6	$(R) - 6$ $(R) - 9$	100 100	80(S) 74 (S)	13	$(R) - 6/(R) - 9$	100	93 (S)			

(the perfluoroarene–methoxyarene  $\pi-\pi$  interactions are possibly too weak in solution), these ligands were tested in the Rh-catalyzed asymmetric hydrogenations of prochiral olefins (dimethyl itaconate and methyl acetamido acrylate) with good results. In two cases, the heterocombination of one methoxybenzyl-derived binaphthyl phosphite  $[(R)-4 \text{ or } (R)-6]$  and the electron-poor pentafluorobenzylderived binaphthyl phosphite  $[(R)-9]$  led to very high enantiomeric excesses, significantly higher than those obtained with the corresponding homocombinations; while both homocomplexes  $(RhL<sub>A</sub>L<sub>A</sub>$  and  $RhL<sub>B</sub>L<sub>B</sub>$ ) were still present (ca. 50%) in the precatalyst mixture, the heterocomplex ( $RhL_A L_B$ , ca. 50%) turned out to be more active and selective, as already reported in many other precedents.[18](#page-4-0)

## Acknowledgements

We thank the European Commission [RTN Network (R)Evolutionary Catalysis MRTN-CT-2006-035866] for financial support and for a predoctoral fellowship (to B. Lynikaite) and a postdoctoral fellowship (to J. Cvengros). C. Gennari gratefully acknowledges Merck Research Laboratories for the Merck's Academic Development Program Award. We also wish to thank Professor Franco Cozzi (Milan University) for reading the manuscript.

## References and notes

- 1. Jerphagnon, T.; Renaud, J.-L.; Bruneau, C. Tetrahedron: Asymmetry 2004, 15, 2101–2111; and references cited therein.
- 2. (a) De Vries, J. G.; Lefort, L. Chem. Eur. J. 2006, 12, 4722–4734; and references cited therein; (b) Jäkel, C.; Paciello, R. Chem. Rev. 2006, 106, 2912–2942.
- 3. Reetz, M. T.; Mehler, G. Angew. Chem., Int. Ed. 2000, 39, 3889–3890.
- 4. (a) Reetz, M. T.; Sell, T. Tetrahedron Lett. 2000, 41, 6333–6336; (b) Claver, C.; Fernandez, E.; Gillon, A.; Heslop, K.; Hyett, D. J.; Martorell, A.; Orpen, A. G.; Pringle, P. G. Chem. Commun. 2000, 961–962.
- 5. van den Berg, M.; Minnaard, A. J.; Schudde, E. P.; Van Esch, J.; de Vries, A. H. M.; de Vries, J. G.; Feringa, B. L. J. Am. Chem. Soc. 2000, 122, 11539–11540.
- 6. (a) Wilkinson, M. J.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Org. Biomol. Chem. 2005, 3, 2371–2383; (b) Breit, B. Angew. Chem., Int. Ed. 2005, 44, 6816–6825; (c) Sandee, A. J.; Reek, J. N. H. Dalton Trans. 2006, 3385–3391.
- 7. (a) Slagt, V. F.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Chem. Commun. 2003, 2474–2475; (b) Slagt, V. F.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Angew. Chem., Int. Ed. 2003, 42, 5619–5623; (c) Slagt, V. F.; Röder, M.; Kamer, P. C. J.; van Leeuwen, P. W. N. M.; Reek, J. N. H. J. Am. Chem. Soc. 2004, 126, 4056–4057; (d) Takacs, J. M.; Reddy, D. S.; Moteki, S. A.; Wu, D.; Palencia, H. J. Am. Chem. Soc. 2004, 126, 4494–4495; (e) Takacs, J. M.; Hrvatin, P. M.; Atkins, J. M.; Reddy, D. S.; Clark, J. L. New J. Chem. 2005, 29, 263–265; (f) Jiang, X.; Lefort, L.; Goudriaan, P. E.; de Vries, A. H. M.; van Leeuwen, P. W. N. M.; de Vries, J. G.; Reek, J. N. H. Angew. Chem., Int. Ed. 2006, 45, 1223–1227; (g) Jiang, X.-B.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Chem. Commun. 2007, 2287–2289; (h) Slagt, V. F.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Dalton Trans. 2007, 2302– 2310.
- 8. (a) Breit, B.; Seiche, W. J. Am. Chem. Soc. 2003, 125, 6608–6609; (b) Breit, B.; Seiche, W. Angew. Chem., Int. Ed. 2005, 44, 1640–1643; (c) Seiche, W.; Schuschkowski, A.; Breit, B. Adv. Synth. Catal. 2005, 347, 1488–1494; (d) Weis, M.; Waloch, C.; Seiche, W.; Breit, B. J. Am. Chem. Soc. 2006, 128, 4188–4189; (e) Chevallier, F.; Breit, B. Angew. Chem., Int. Ed. 2006, 45, 1599–1602; (f) Knight, L. K.; Freixa, Z.; van Leeuwen, P. W. N. M.; Reek, J. N. H. Organometallics 2006, 24, 954– 960; (g) Hattori, G.; Hori, T.; Miyake, Y.; Nishibayashi, Y. J. Am. Chem. Soc. 2007, 129, 12930–12931; (h) Sandee, A. J.; van der Burg, A. M.; Reek, J. N. H. Chem. Commun. 2007, 864–866; (i) Waloch, C.; Wieland, J.; Keller, M.; Breit, B. Angew. Chem., Int. Ed. 2007, 46, 3037–3039; (j) Birkholz, M.-N.; Dubrovina, N. V.; Jiao, H.; Michalik, D.; Holz, J.; Paciello, R.; Breit, B.; Börner, A. Chem. Eur. J. 2007, 13, 5896–5907; (k) Birkholz, M.-N.; Dubrovina, N. V.; Shuklov, I. A.; Holz, J.; Paciello, R.; Walloch, C.; Breit, B.; Börner, A. Tetrahedron: Asymmetry 2007, 18, 2055–2060.
- 9. Gulyàs, H.; Benet-Buchholz, J.; Escudero-Adan, E. C.; Freixa, Z.; van Leeuwen, P. W. N. M. Chem. Eur. J. 2007, 13, 3424–3430.
- 10. Machut, C.; Patrigeon, J.; Tilloy, S.; Bricout, H.; Hapiot, F.; Monflier, E. Angew. Chem., Int. Ed. 2007, 46, 3040–3042.
- 11. (a) Patrick, C. R.; Prosser, G. S. Nature 1960, 187, 1021; For comprehensive reviews, see: (b) Williams, J. H. Acc. Chem. Res. 1993, 26, 593–598; (c) Meyer, E. A.; Castellano, R. K.; Diederich, F. Angew. Chem., Int. Ed. 2003, 42, 1210–1250.
- 12. For recent examples, see: (a) Gdaniec, M.; Jankowski, W.; Milewska, M. J.; Polonski, T. Angew. Chem., Int. Ed. 2003, 42, 3903–3906; (b) Facchetti, A.; Yoon, M.-H.; Stern, C. L.; Katz, H. E.; Marks, T. J. Angew. Chem., Int. Ed. 2003, 42, 3900–3903; (c) Watt, S. W.; Dai, C.; Scott, A. J.; Burke, J. M.; Thomas, R. L.; Collings, J. C.; Viney, C.; Clegg, W.; Marder, T. B. Angew. Chem., Int. Ed. 2004, 43, 3061–3063; (d) Knapp, C.; Lork, E.; Mews, R.; Zibarev, A. V. Eur. J. Inorg. Chem. 2004, 12, 2446–2451; (e) Bacchi, S.; Benaglia, M.; Cozzi, F.; Demartin, F.; Filippini, G.; Gavezzotti, A. Chem. Eur. J. 2006, 12, 3538–3546; (f) Cozzi, F.; Bacchi, S.; Filippini, G.; Pilati, T.; Gavezzotti, A. Chem. Eur. J. 2007, 13, 7177–7184; (g) Hori, A.; Shinohe, A.; Yamasaki, M.; Nishibori, E.; Aoyagi, S.; Sakata, M. Angew. Chem., Int. Ed. 2007, 46, 7617–7620.
- 13. (a) Weck, M.; Dunn, A. R.; Matsumoto, K.; Coates, G. W.; Lobkovsky, E. B.; Grubbs, R. H. Angew. Chem., Int. Ed. 1999, 38, 2741–2745; (b) Ponzini, F.; Zagha, R.; Hardcastle, K.; Siegel, J. S. Angew. Chem., Int. Ed. 2000, 39, 2323–2325; (c) Collings, J. C.; Roscoe, K. P.; Thomas, R. L.; Batsanov, A. S.; Stimson, L. M.; Howard, J. A. K.; Marder, T. B. New J. Chem. 2002, 26, 1740–1746; (d) Reichenbaecher, K.; Suess, H. I.; Hulliger, J. Chem. Soc. Rev. 2005, 34, 22–30.
- 14. Coates, G. W.; Dunn, A. R.; Henning, L. M.; Ziller, J. W.; Lobkovsky, E. B.; Grubbs, R. H. J. Am. Chem. Soc. 1998, 120, 3641–3649.
- 15. (a) Cozzi, F.; Ponzini, F.; Annunziata, R.; Cinquini, M.; Siegel, J. S. Angew. Chem., Int. Ed Engl. 1995, 34, 1019–1020; (b) Korenaga, T.;

<span id="page-4-0"></span>Kawauchi, Y.; Kosaki, T.; Ema, T.; Sakai, T. Bull. Chem. Soc. Jpn. 2005, 78, 2175-2179; (c) Pérez-Casa, S.; Hernández-Trujillo, J.; Costas, M. J. Phys. Chem. 2003, 107, 4167–4174.

- 16. (a) Marsella, M. J.; Wang, Z.-Q.; Reid, R. J.; Yoon, K. Org. Lett. 2001, 3, 885–887; (b) El-azizi, Y.; Schmitzer, A.; Collins, S. K. Angew. Chem., Int. Ed. 2006, 45, 968–973; (c) Collins, S. K.; El-Azizi, Y. Pure Appl. Chem. 2006, 78, 783–789; (d) Collins, S. K.; El-azizi, Y.; Schmitzer, A. J. Org. Chem. 2007, 72, 6397–6408.
- 17. Franciò, G.; Arena, C. G.; Faraone, F.; Graiff, C.; Lanfranchi, M.; Tiripicchio, A. Eur. J. Inorg. Chem. 1999, 1219–1227.
- 18. (a) Reetz, M. T.; Sell, T.; Meiswinkel, A.; Mehler, G. Patent application DE-A 10247633.0 (11-10-2002); (b) Reetz, M. T.; Sell, T.; Meiswinkel, A.; Mehler, G. Angew. Chem., Int. Ed. 2003, 42, 790– 793; (c) Peña, D.; Minnaard, A. J.; Boogers, J. A. F.; de Vries, A. H. M.; de Vries, J. G.; Feringa, B. L. Org. Biomol. Chem. 2003, 1, 1087– 1089; (d) Duursma, A.; Hoen, R.; Schuppan, J.; Hulst, R.; Minnaard, A. J.; Feringa, B. L. Org. Lett. 2003, 5, 3111–3113; (e) Reetz, M. T.; Mehler, G. Tetrahedron Lett. 2003, 44, 4593–4596; (f) Monti, C.;

Gennari, C.; Piarulli, U. Tetrahedron Lett. 2004, 45, 6859–6862; (g) Reetz, M. T.; Mehler, G.; Meiswinkel, A. Tetrahedron: Asymmetry 2004, 15, 2165–2167; (h) Reetz, M. T.; Li, X. Tetrahedron 2004, 60, 9709–9714; (i) Reetz, M. T.; Li, X. Angew. Chem., Int. Ed. 2005, 44, 2959–2962; (j) Reetz, M. T.; Li, X. Angew. Chem., Int. Ed. 2005, 44, 2962–2964; (k) Monti, C.; Gennari, C.; Piarulli, U. Chem. Commun. 2005, 5281–5283; (l) Monti, C.; Gennari, C.; Piarulli, U.; De Vries, J. G.; De Vries, A. H. M.; Lefort, L. Chem. Eur. J. 2005, 11, 6701–6717; (m) Hoen, R.; Boogers, J. A. F.; Bernsmann, H.; Minnaard, A. J.; Meetsma, A.; Tiemersma-Wegman, T. D.; de Vries, A. H. M.; de Vries, J. G.; Feringa, B. L. Angew. Chem., Int. Ed. 2005, 44, 4209– 4212; (n) Duursma, A.; Peña, D.; Minnaard, A. J.; Feringa, B. L. Tetrahedron: Asymmetry 2005, 16, 1901–1904; (o) Reetz, M. T.; Fu, Y.; Meiswinkel, A. Angew. Chem., Int. Ed. 2006, 45, 1412–1415; (p) Gennari, C.; Monti, C.; Piarulli, U. Pure Appl. Chem. 2006, 78, 303– 310; (q) Monti, C.; Gennari, C.; Piarulli, U. Chem. Eur. J. 2007, 13, 1547–1558; (r) Marelli, C.; Monti, C.; Gennari, C.; Piarulli, U. Synlett 2007, 2213–2216.